

9/6/0

(12) **UK Patent Application** (19) **GB** (11) **2 240 284** (13) **A**
(43) Date of A publication 31.07.1991

(21) Application No 9100951.4

(22) Date of filing 16.01.1991

(30) Priority data

(31) 19162A90

(32) 26.01.1990

(33) IT

(71) Applicant

Snamprogetti S.p.A.

(Incorporated in Italy)

Corso Venezia 16, Milan, Italy

(72) Inventors

Luca Basini

Mario Marchionna

Stefano Rossini

Domenico Sanfilippo

(74) Agent and/or Address for Service

Haseltine Lake & Co

Hazlitt House, 28 Southampton Buildings,

Chancery Lane, London, WC2A 1AT, United Kingdom

(51) INT CL⁵

B01J 23/40, C01B 3/40

(52) UK CL (Edition K)

B1E EAA E1119 E1180 E1206 E1285 E1298 E1315

E1342 E1471 E1472 E1496 E1513 E1634 E1701

E1732 E1738

C5E EAR E101 E111 E122 E124 E126

U1S S1360 S1483 S1502

(56) Documents cited

GB 1600587 A

EP 0206535 A1

EP 0105663 A2

EP 0084273 A2

US 4581157 A

(58) Field of search

UK CL (Edition K) B1E

INT CL⁵ B01J, C01B

ONLINE: WPI AND CLAIMS

(54) **Catalytic system and process for producing synthesis gas by reforming light hydrocarbons with CO₂**

(57) A catalytic system for the production of synthesis gas by reacting light hydrocarbons, preferably methane, with CO₂ is described formed from:

- one or more compounds of metals of the platinum group, preferably chosen from rhodium, ruthenium and iridium;
- a support consisting of inorganic compounds chosen from oxides and/or spinels of aluminium, magnesium, zirconium, silicon, cerium and/or lanthanum, possibly in the presence of alkaline metals,

in which the weight percentage of the metal or metals of the platinum group in the catalytic system is between 0.01 and 20%, and preferably between 0.1 and 5%.

GB 2 240 284 A

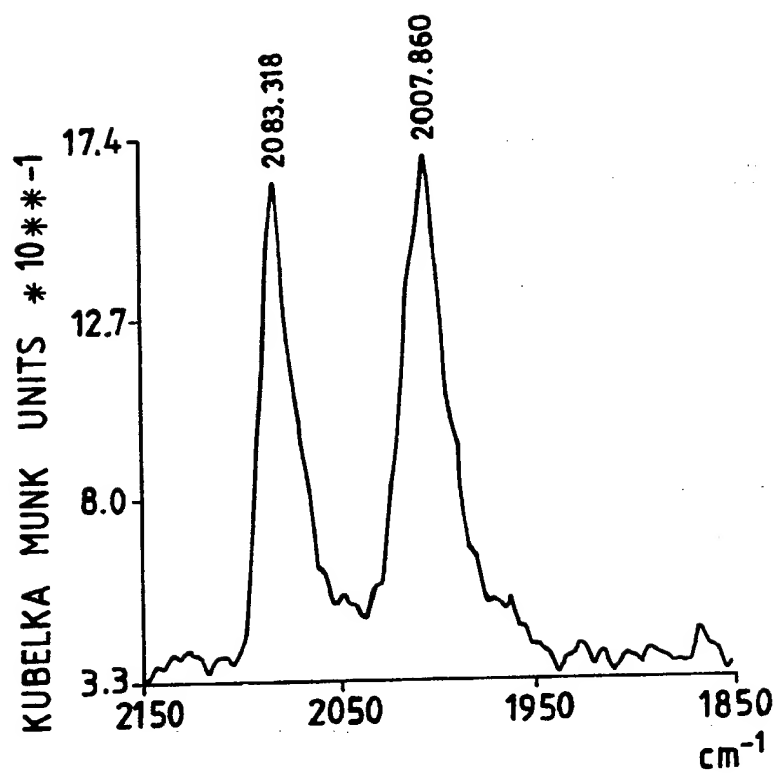
Fig.1

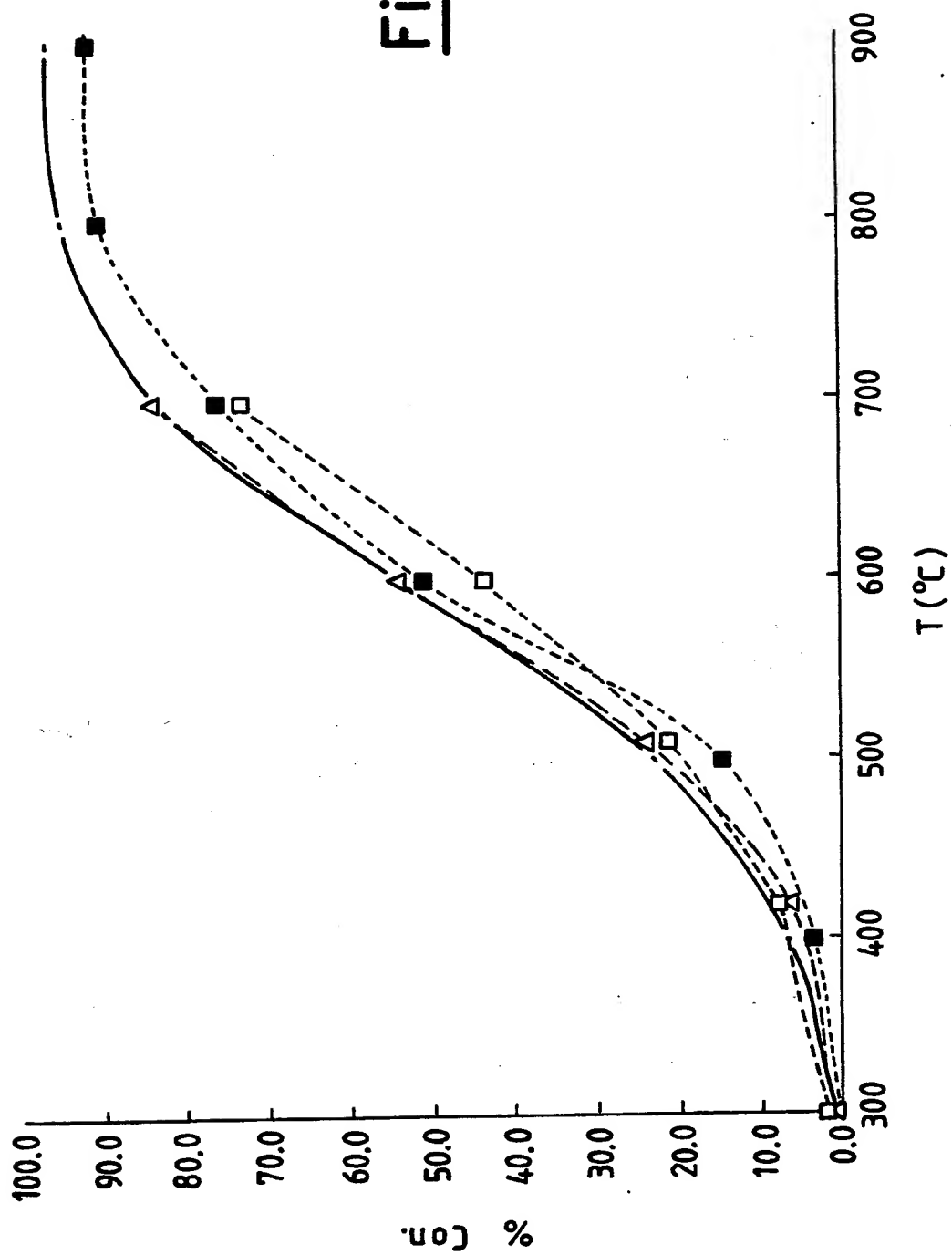
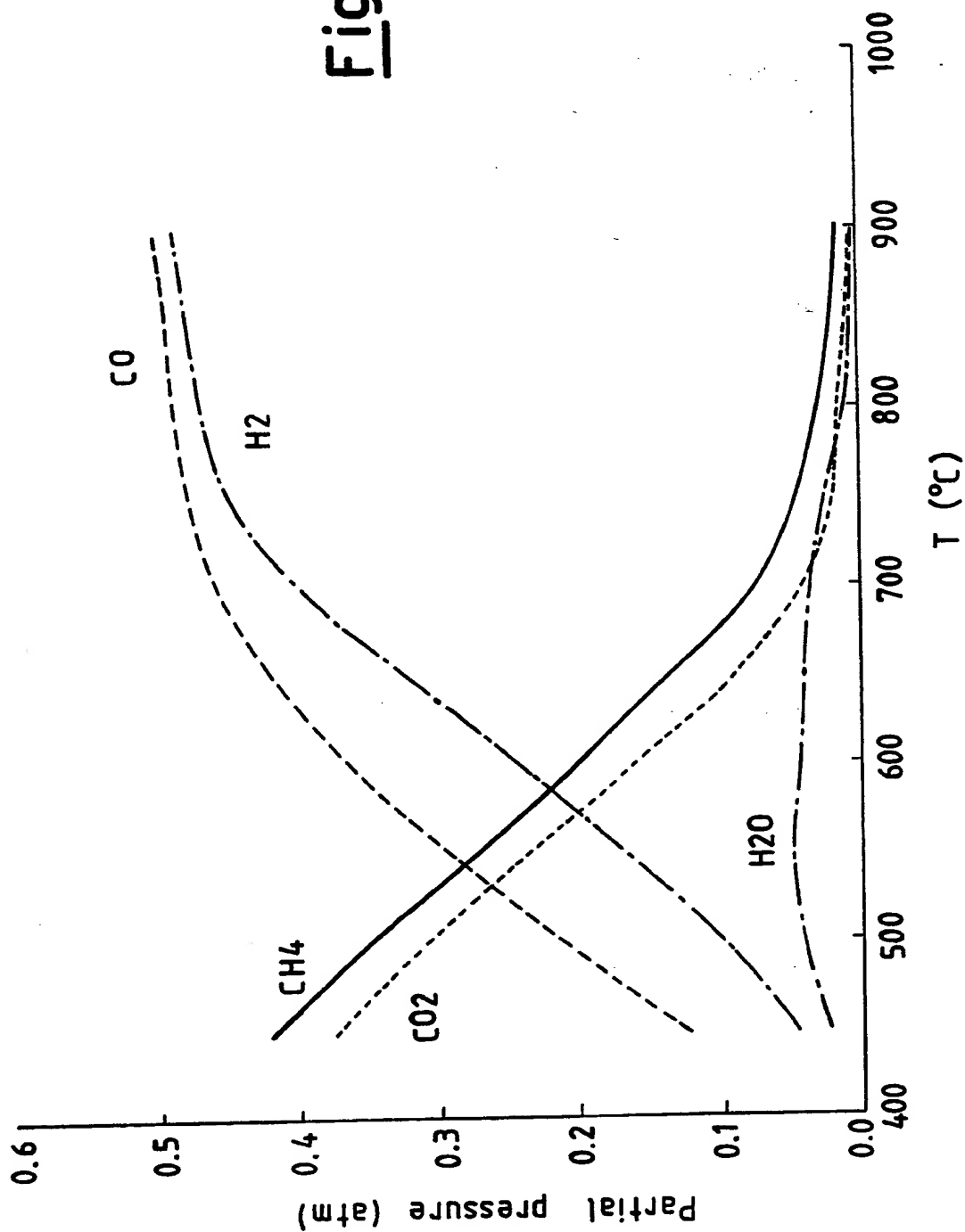
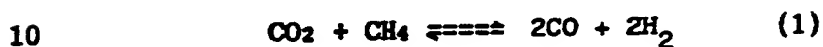
Fig.2

Fig.3

CATALYTIC SYSTEM AND PROCESS FOR PRODUCING SYNTHESIS GAS BY
REFORMING LIGHT HYDROCARBONS WITH CO₂

This invention relates to a catalytic system and its use in a
5 reforming process for the single-stage production of a gaseous
mixture of H₂ and CO.

The main reactants used are CO₂ and light hydrocarbons, preferably
methane. The chemical equation describing the process of the
present invention is:



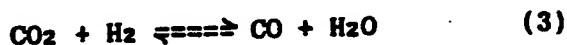
Hydrocarbon reforming reactions using CO₂ have certain
considerable advantages compared with the widespread steam
reforming processes described by the chemical equation:



15 Processes mainly using reaction (1) represent the best method for
producing mixtures of H₂ and CO if the natural gas used as the
feedstock contains large CO₂ quantities.

In addition an H₂/CO mixture in a ratio close to 1, as can be
easily obtained by this invention, can be used advantageously in
20 alcohol synthesis and in oxosynthesis. Currently, using syngas
produced by the steam reforming reaction (2), the obtained H₂/CO
mixture has a ratio ≥ 3 . To obtain smaller ratios a second stage

has to be used employing the reaction:



Adjusting the CO/H₂ ratio by this chemical reaction negatively affects the overall economy of the process.

- 5 A potential use of a reforming process using CO₂ as the main reactant instead of steam is in Fischer-Tropsch synthesis plants in which the CO₂ and methane produced could be again recycled to syngas of low H₂/CO ratio.

- 10 The reforming process described by reaction (1), which produces H₂/CO mixtures in an approximately equimolecular ratio, in a single stage, can also be advantageously used in highly integrated plants for ferrous mineral reduction.

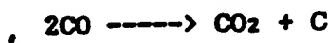
- Finally, methane reforming processes employing CO₂ can be advantageously used, compared with steam reforming reactions, in
15 thermal cycles for energy storage and transport by thermochemical pipe (TCP) [see T.A. Chubb, Solar Energy, 24, (1980) 341].

- However, in contrast to steam reforming, systems involving H₂ and CO synthesis from CO₂ and from light hydrocarbons do not have a well defined technology behind them. The Ni-based catalysts usually used in
20 steam reforming processes are not sufficiently selective, and deactivate rapidly when the H₂O/C ratio is less than 2 [see R.E. Reitmeier, K. Atwood, H.A. Bennet Jr. and H.M. Baugh, Ind. Eng. Chem. 40 (4), 620 (1948)].

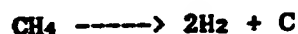
- This deactivation is due to the formation of carbon, which covers
25 the active metal centres during catalysis and accumulates in the catalyst pores, possibly causing fragmentation.

A catalytic system has now been found which produces synthesis gas

(H₂ and CO) by a light hydrocarbon reforming reaction without undergoing any discernible deactivation due to coke formation by the reactions:



5



even if the H₂O/C ratio is distinctly favourable to such formation.

The catalytic system according to the present invention is characterised by being formed from:

- 10 - one or more compounds of metals of the platinum group, preferably chosen from rhodium, ruthenium and iridium;
- a support consisting of inorganic compounds chosen from oxides and/or spinels of aluminium, magnesium, zirconium, silicon, cerium and/or lanthanum, either alone or in mutual combination
15 and possibly in the presence of alkaline metals,

in which the weight percentage of the metal or metals of the platinum group in the catalytic system is between 0.01 and 20%, and preferably between 0.1 and 5%.

The supports used can also consist of silicified aluminium,

- 20 magnesium, cerium or lanthanum oxides.

The surface area of the catalysts used preferably varies between 1 and 400 m²/g and more preferably between 10 and 200 m²/g, while the pore volume preferably varies between 0.1 and 3 cc/g and more preferably between 0.5 and 2 cc/g.

- 25 The catalytic system can be obtained either by impregnating the inorganic compounds with a solution of a salt of the metals of the platinum group followed by thermal drying and calcining, or by

dispersing the inorganic compounds in an organic solvent, then reacting this in a carbon monoxide or inert atmosphere with solutions of compounds of the metals of the platinum group.

In this second case, the generally exothermic reaction, which
5 results in coloured reaction products, is followed by filtration, drying and calcining.

More particularly, the catalytic system in question can be prepared by heterogeneous solid-liquid reaction at a temperature of between 0°C and 150°C, and preferably between 20°C and 50°C,
10 between compounds of the metals of the platinum group dissolved in an organic solvent and the stated inorganic compounds dispersed in the same solvent.

Following this procedure, the metal quantity which fixes to the substrate is determined mainly by the chemical properties of the
15 inorganic oxide rather than by its porosity and surface area.

These latter are however important with respect to the integrity and stability of the catalyst during the reforming reaction. In this respect, the accumulation of carbon in too small pores leads to material fragmentation. A reduced support surface area also
20 results in a lesser dispersion of metal and favours sinter phenomena with consequent catalyst deactivation.

The method for preparing supports consisting of silicified aluminium, magnesium, cerium or lanthanum oxides consists essentially of a condensation reaction between the inorganic oxide
25 (of aluminium, magnesium, cerium or lanthanum) and a silicon compound containing hydrolyzable organic groups, followed by removal of the unhydrolyzed organic residues by combustion or

reaction in the presence of steam.

Using such silicification methods, materials can be obtained containing percentages of silicon varying between 0.5 and 15% and preferably between 1 and 5% by weight.

- 5 The present invention also provides a catalytic reforming process for light hydrocarbons, preferably methane, which enables mixtures of H_2 and CO to be obtained in ^{molecular} / ratios varying between 0.6 and 6, and preferably between 0.8 and 3.

This catalytic process is characterised by conducting the
10 reforming preferably in a single stage using the aforescribed catalytic system and operating at a temperature of between 350°C and 850°C, and preferably between 550°C and 750°C, at a pressure of between 0.5 and 50 atm, and preferably between 1 and 40 atm.
If methane is used, the required volumetric CO_2/CH_4 reactant ratio
15 is between 0.5 and 15, and preferably between 0.8 and 10.

Under all these thermodynamic conditions the process can also be conducted in the presence of steam, if the particular application of the product synthesis gas requires it.

In this respect, it is necessary only to adjust the relative CO_2
20 and H_2O feed quantities to obtain a synthesis gas with any desired H_2/CO ratio from 1 to 6.

Although the process is particularly suitable for methane reforming reactions, any other light hydrocarbon or mixture can be used in the process.

- 25 For example C_1-C_4 paraffins and olefins can be used by suitably choosing optimum temperature and pressure conditions and CO_2 ratios.

Any natural gas containing hydrocarbon mixtures in which the methane content preferably exceeds 80% by volume can be used. Some examples are given below to better illustrate the invention, which however is not to be considered limited by them or to them.

5 **EXAMPLE 1**

Catalyst preparation

The inorganic oxide used as the support was prepared by the following procedure.

A commercial magnesium oxide (supplied by Carlo Erba) with a surface area of 210 m²/g was suspended under stirring in a tetraethylsilicate (TES) solution. The temperature was maintained between 80°C and 90°C to favour evaporation of the ethanol formed by the condensation reactions. A dry gaseous nitrogen stream was fed into the reaction environment. Gas chromatograph analysis of the exit gas showed the formation of ethanol.

The condensation reaction was considered at an end when ethanol was no longer detected in the exit gas stream. At this point the temperature was raised to 180°C to distil off the unreacted TES. The unreacted ethoxy groups bonded to silicon atoms anchored to the solid inorganic support were then hydrolyzed by feeding a stream of nitrogen and steam at 200°C. Ethanol was also detected in the gas stream during this step. Infrared spectrum analysis on the material obtained up to this point shows the presence of numerous hydroxyl bands which were not present in the starting material. The solid was then heated to 850°C (5°C/min) and maintained at this temperature for 10 hours. After this treatment the surface area had reduced to 32m²/g, the silicon content being

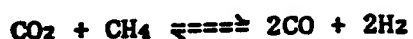
1.5% by weight. Differential thermal, thermogravimetric and infrared spectroscopic analysis conducted during three cycles at temperatures between 25°C and 750°C showed no significant alteration in the physico-chemical properties of the silicified materials obtained. 50 g of silicified magnesium oxide were then suspended in 100 ml of 2-methylpentane in a nitrogen atmosphere. A second solution of 50 ml of the same solvent containing 0.91 g of $\text{Rh}_4(\text{CO})_{12}$ in a CO atmosphere was dripped rapidly into the silicified oxide suspension under stirring. The organic solution decolours rapidly passing from intense red to colourless, with the white solid simultaneously colouring. It is filtered in an inert atmosphere to obtain a material containing 1% by weight of Rh in a highly dispersed condition, as could be deduced from an analysis of the vibrational carbonyl bands of the surface complexes (see Figure 1 showing the diffused reflectance spectrum obtained on the pulverulent solid, in which the horizontal axis represents the wave number in cm^{-1} and the vertical axis represents Kubelka Munk intensity units). The transformations of the surface complexes during thermal reduction with hydrogen in gaseous CH_4 and CO_2 atmospheres were also studied by infrared spectroscopy. This resulted in a satisfactory understanding of surface nucleation phenomena, ensuring high reproducibility in material preparation.

Reforming reaction

The reforming reaction was conducted in a fixed bed quartz reactor containing 3 cc of catalyst by feeding a gaseous equimolecular stream of CH_4 and CO_2 at a pressure of one atmosphere. In-line gas chromatograph analysis was carried out on the exit gas stream

starting from 300°C and continuing until 750°C. The gas hourly space velocity was maintained at 1000 (l/kg.h).

Figure 2 shows the various experimental CH₄ and CO₂ conversion values at the various temperatures investigated (shown by black squares and dots respectively). The same figure also shows the theoretically calculated conversion values for the equilibrium system for the reactions:



(A)

10



The theoretical conversion of CO₂ at equilibrium is shown by triangles, and that of CH₄ by white squares.

From the results obtained it can be deduced that the catalyst is extremely active and enables conversions close to the thermodynamic equilibrium conversions to be obtained within the temperature range studied. The H₂O percentage in the reaction product mixture is also close to the values calculated for the system (A) at equilibrium. Figure 3 shows the theoretically calculated variations in the concentrations of the gaseous species with temperature for the system under examination at a total pressure of 1 atm. The experimentally obtained gaseous species concentrations faithfully reproduce this pattern.

The H₂/CO ratio within the range of 650-750°C was slightly less than 1.

Table 1 shows the results obtained for catalytic tests lasting 100 hours at 700°C conducted with the catalysts described in Examples 1-4 and 6, compared with the results obtained using a commercial

steam reforming catalyst (Example 7) containing approximately 15.5% of Ni supported on α -alumina.

During the tests, catalysts comprising Rh deposited on silicified magnesium oxide proved to be extremely active in catalyzing the reactions of the system (A) but, surprisingly, not in catalyzing the carbon formation reactions even where these are favoured under these conditions.

The results of quantitative carbon analysis on the discharged catalysts are shown in Table 1. During the 100 hours the activity and selectivity of the catalytic system remained constant.

EXAMPLE 2

The catalyst synthesis procedure described in Example 1 was repeated but using a solution containing 1.05 g of $\text{Ru}(\text{CO})_{12}$ to obtain a solid containing 1% of Ru by weight.

The reforming reaction was conducted as in Example 1 feeding the same reactant mixture under the same pressure and spatial velocity conditions at temperatures of between 300°C and 750°C. Again in this case the CH_4 and CO_2 conversions are close to equilibrium values even if slightly less than those obtained in Example 1 (see Table 1).

EXAMPLES 3-4

In these cases the catalysts used contained the noble metals Rh (Example 3) or Ru (Example 4) and silicified alumina. This latter was prepared by condensing tetraethylsilicate with a gamma alumina supplied by AKZO in accordance with the procedure described in Example 1. The catalytic systems thus obtained proved to possess the same characteristics as those described in Examples 1 and 2,

ie active in catalyzing the reactions of the system (A) within the entire temperature range investigated but inactive in catalyzing the reactions involved in the formation of carbon on the catalyst. Table 1 also shows the results obtained during 100 hour catalytic tests in these two cases.

EXAMPLE 5

In this case the reactants were a gaseous stream of C_2H_6 and CO_2 in a 1/2 ratio. The catalytic tests were conducted at temperatures between $400^\circ C$ and $700^\circ C$ using the catalyst of Example 2. The ethane conversion/during a catalytic test lasting 100 hours was found to be 100% and the CO_2 conversion 98%. The H_2/CO ratio was 0.7. The methane in the exit gas was less than 3%.

EXAMPLE 6

In this example the catalyst synthesis procedure described in Example 1 was modified in that the noble metal was deposited on the silicified oxide by an impregnation reaction conducted by dripping an aqueous solution of Rh nitrate onto the silicified oxide until it was just soaked. The catalyst obtained in this manner contained 1% (wt/wt) of Rh. In this case the catalytic tests conducted as in Examples 1-5 showed that the material modifies its characteristics during the first 10 hours of reaction at $700^\circ C$. During this period the conversion values increase until they settle down at the values shown in Table 1. However after the induction period the characteristics of the catalytic systems described in Examples 1-4 are again obtained in this case.

EXAMPLE 7 - Comparative

Compared with Example 1 a commercial steam reforming catalyst consisting of about 15.5% by weight of Ni supported on alumina was used.

The results obtained are shown in Table 1.

5

TABLE 1

	%conv CO ₂	% conv CH ₄	sel.	H ₂ /CO	mg C/g cat
			*	(mol/mol)	(100 h)
Ex.1	84.1	73.4	96.2	0.87	< 0.5
10 Ex.2	75.6	68.3	97.1	0.90	0.5
Ex.3	81.4	70.1	96.0	0.85	< 0.5
Ex.4	71.5	63.7	96.7	0.88	0.8
Ex.6	81.2	70.8	96.0	0.85	0.7
Ex.7	68.5	64.3	85.2	0.92	65.2

15 * (moles CO + H₂)/(moles CO + H₂ + H₂O + C).100

CLAIMS

1. A catalytic system for the production of synthesis gas by reforming hydrocarbon with CO_2 , comprising:

5 one or more compounds of a platinum group metal; and

a support comprising an inorganic compound chosen from oxides and/or spinels of aluminium, magnesium, zirconium, silicon, cerium and/or lanthanum, either
10 alone or in combination;

in which the weight of the platinum group metal or metals in the catalytic system is from 0.01 to 20%.

2. A catalytic system as claimed in claim 1, wherein the weight percentage of the platinum group
15 metal or metals is from 0.1 to 5%.

3. A catalytic system as claimed in claim 1 or 2, wherein the platinum group metal is rhodium, ruthenium or iridium.

4. A catalytic system as claimed in any of
20 claims 1 to 3, having a surface area of from 1 to 400 m^2/g and a pore volume of from 0.1 to 3 cc/g .

5. A catalytic system as claimed in claim 4, having a surface area of from 10 to 200 m^2/g and a pore volume from 0.5 to 2 cc/g .

25 6. A catalytic system as claimed in any of claims 1 to 5, wherein the support comprises silicified aluminium, magnesium, cerium or lanthanum oxide.

7. A catalytic system as claimed in any of claims 1 to 6, further comprising one or more alkali
30 metals or alkaline earth metals on the support.

8. A catalytic system as claimed in claim 12, substantially as hereinbefore described.

9. A process for producing synthesis gas by reforming hydrocarbon with CO_2 , wherein the reforming is
35 carried out in the presence of a catalytic system as claimed in any of claims 1 to 8, at a temperature of

from 350°C to 850°C and at a pressure of from 0.5 to 50 atm.

10. A process as claimed in claim 9, wherein the temperature is from 550 to 750°C and the pressure is
5 from 1 to 40 atm.

11. A process as claimed in claim 9 or 10, wherein the reforming is conducted in a single stage.

12. A process as claimed in any of claims 9 to 11, wherein the volumetric CO₂/hydrocarbon ratio is from
10 0.5 to 15.

13. A process as claimed in claim 12, wherein the volumetric CO₂/hydrocarbon ratio is from 0.8 to 10.

14. A process as claimed in any of claims 9 to 12, wherein steam is present to such an extent that the
15 H₂/CO ratio is from 1 to 6.

15. A process as claimed in claim 9, wherein the hydrocarbon is methane.

16. A process as claimed in claim 9, substantially as hereinbefore described.

20 17. Synthesis gas produced by a process as claimed in any of claims 9 to 16.